

ALCOHOL, PHENOL & ETHER

ALCOHOL

Hydroxy derivatives

Aliphatic hydroxy derivatives

Aromatic hydroxy derivatives

(I) Aliphatic hydroxy derivatives :

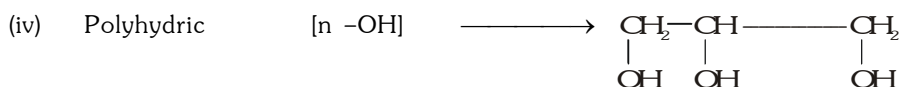
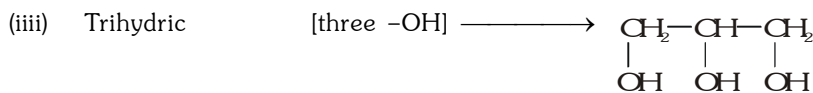
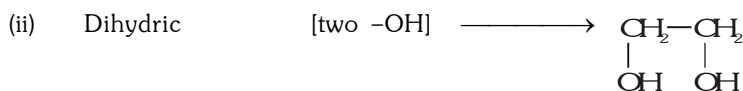
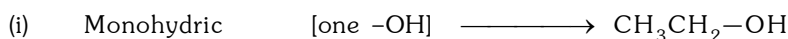
Hydroxy derivatives in which -OH is directly attached to sp^3 C (Alcoholic compounds).

(II) Aromatic hydroxy derivatives :

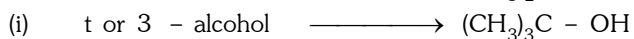
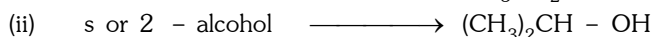
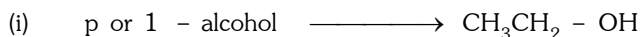
Hydroxy derivatives in which -OH is directly attached to sp^2 C or benzene ring (Phenolic compounds).

□ Aliphatic hydroxy derivatives :

(a) Classification according to number of -OH groups :

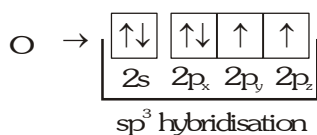
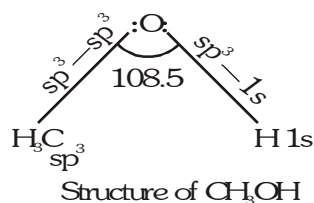


(b) Classification according to nature of carbon :



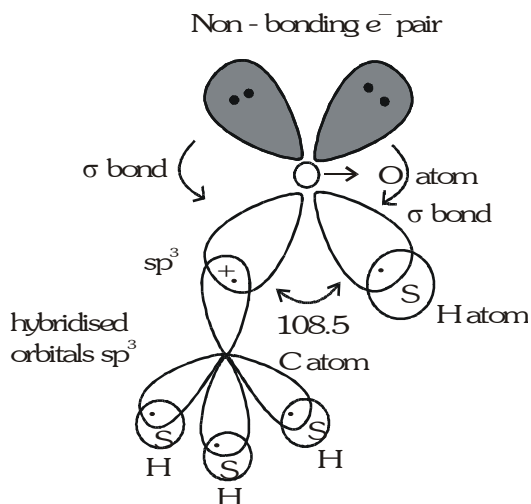
□ Structure of alcohol :

Alcohols are bent molecules. The carbon atom (linked with 'O' atom of -OH group) is sp^3 hybridised. The central 'O' atom is also in sp^3 state of hybridisation. The bond angle is 108.5° . In sp^3 hybridisation of O - $2s^2, 2p_x^2, 2p_y^1, 2p_z^1$ orbitals hybridised to form sp^3 orbitals



In these four orbitals two containing one electron each and two containing two electrons each. Orbitals containing two electrons do not take part in bonding. Other two half filled orbitals form σ bond with s-orbitals of H -atom and hybridised orbital of C-atom (O-C).

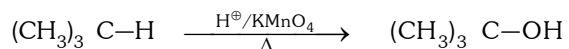
Due to lone pair effect the bond angle of tetrahedral oxygen atom is lesser than normal tetrahedral structure ($109^\circ 28'$).



MONOHYDRIC ALCOHOL

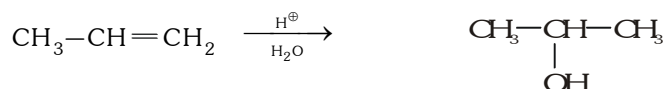
□ **General methods of preparation :**

(a) **From alkanes (By oxidation) :**



(b) **From alkenes :**

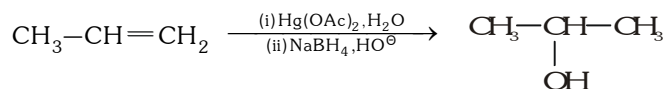
(i) **By hydration :**



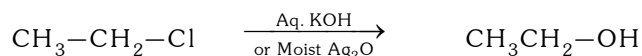
(ii) **By hydroboration oxidation :**



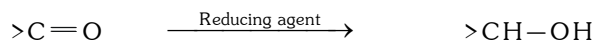
(iii) **By oxymercuration demercuration :**



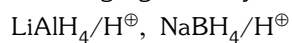
(c) **From alkyl halides (By hydrolysis) :**



(d) **From carbonyl compounds (By reduction) :**



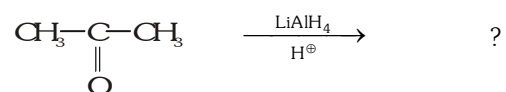
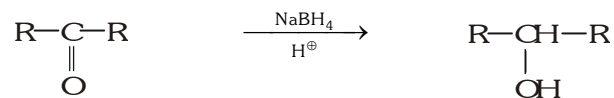
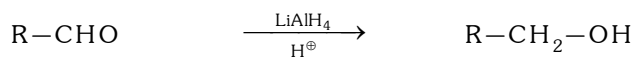
◆ Reducing agents may be,



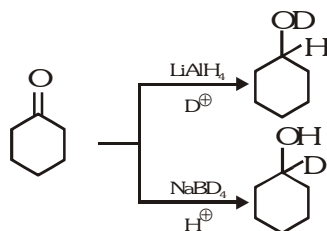
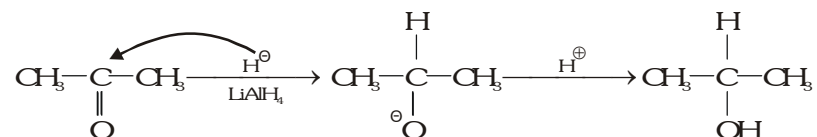
$\text{Na} + \text{EtOH}$ [Bouveault-blanc Reduction]

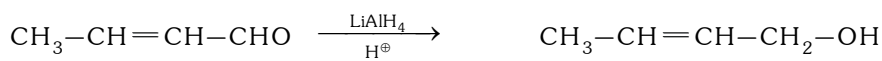
NaH [Darzen reduction]

Ni/H_2

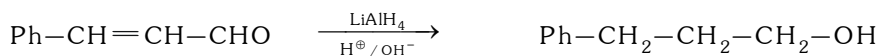


◆ **Mechanism :**



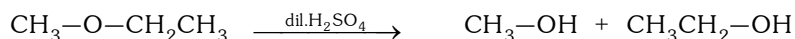
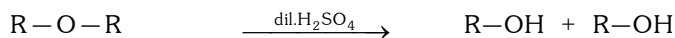


Crotonaldehyde

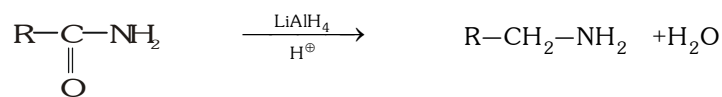
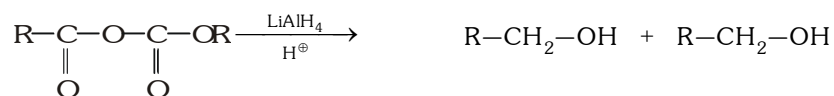
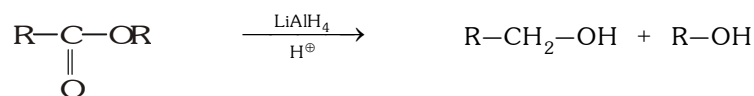
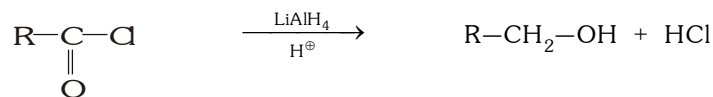
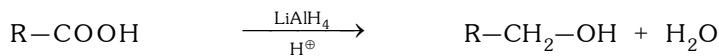


Cinnamaldehyde

(e) From ethers :

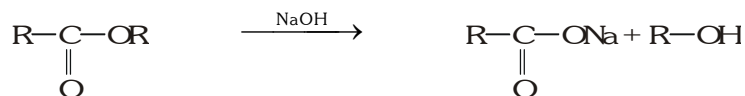


(f) From acid and derivatives (By reduction) :

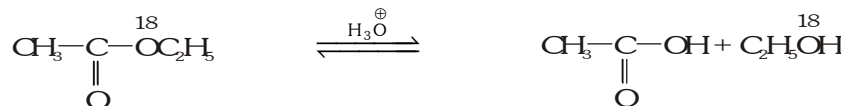
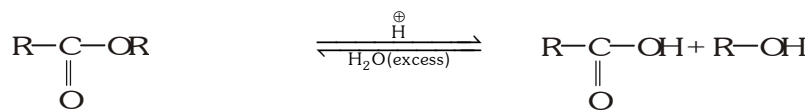


(g) From esters (By hydrolysis) :

(i) By alkaline hydrolysis :

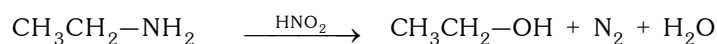
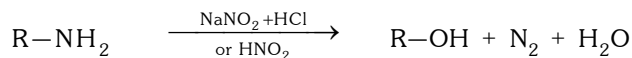


(ii) By acidic hydrolysis :

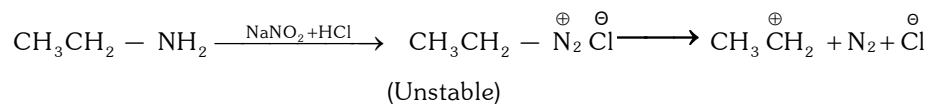


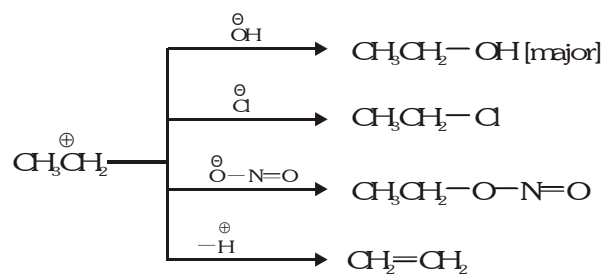
This reaction is reversible reaction and its order is 1 and it is also called Pseudo-Unimolecular reaction.

(h) From p-amines :

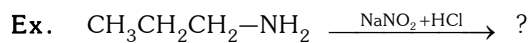


◆ Mechanism :

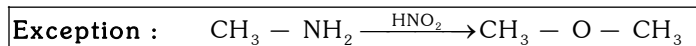
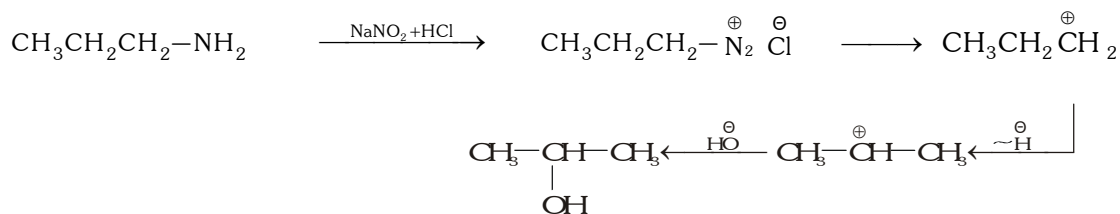




Intermediate is carbocation so rearrangement may be possible.

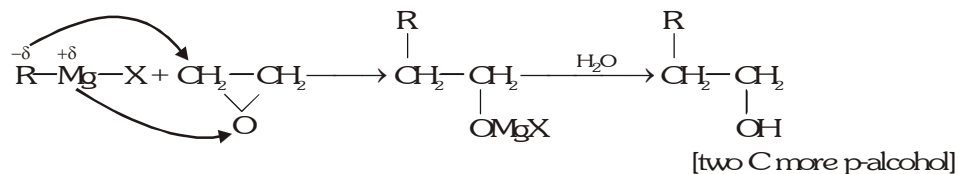
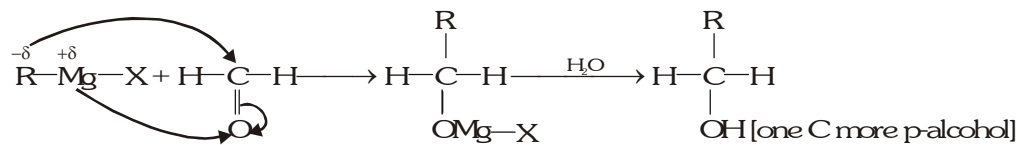
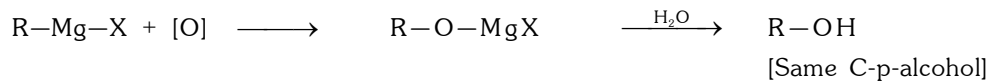


Sol. Mechanism :

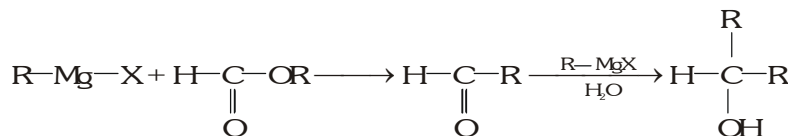
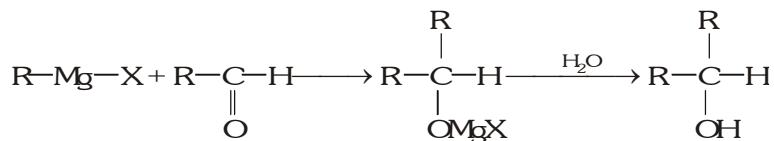


(i) From Grignard reagent :

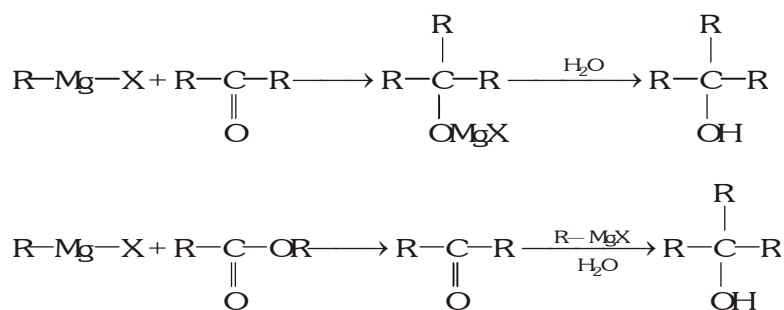
(i) p-alcohol :



(ii) s-alcohol :



(iii) t-alcohol :

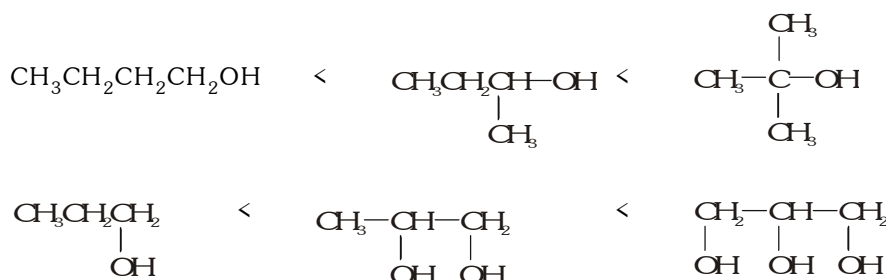
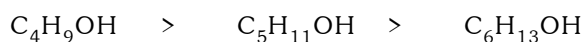


□ **Physical properties :**

- (i) C_1 to C_{11} are colourless liquids and high alcohols are solids.
- (ii) Density of monohydric alcohol is less than H_2O .
- (iii) Density \propto mol. wt. (for monohydric alcohol).
- (iv) **Solubility** : C_1 to C_3 and t-butyl alcohol is completely soluble in H_2O due to H-bonding.

$$\text{solubility} \propto \text{No. of side chains} \propto \frac{1}{\text{molecular weight}}$$

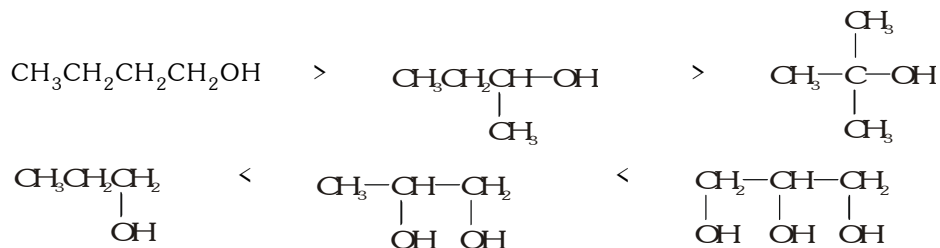
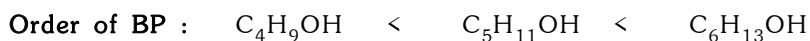
Order of solubility :



[Number of $-\text{OH}$ increases, H-bonding increases]

- (v) **Boiling points** : B.P. \propto molecular weight

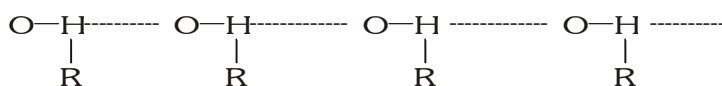
$$\text{If molecular wt. is same then B.P.} \propto \frac{1}{\text{branching}}$$



[Number of OH increases, H-bonding increases]

Ex. Boiling point of alcohol is more than corresponding ether. Why ?

Sol. Reason : H-bonding in alcohol.



Ex. Boiling point of alcohol is less than corresponding carboxylic acid. Why ?

Sol. Reason : Dimer formation in carboxylic acid.



□ **Chemical properties :**

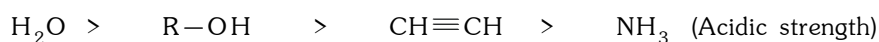
Monohydric alcohol show following reactions

- (A) Reaction involving cleavage of $\text{O}-\text{H}$
- (B) Reaction involving cleavage of $\text{C}-\text{OH}$
- (C) Reaction involving complete molecule of alcohol

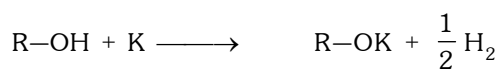
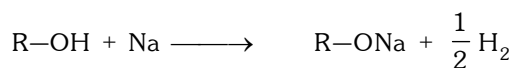
(A) Reaction involving cleavage of $\text{O}-\text{H}$: Reactivity order (Acidic nature) is



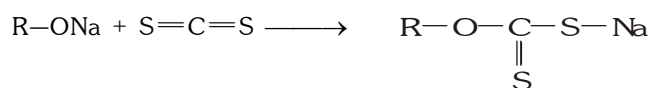
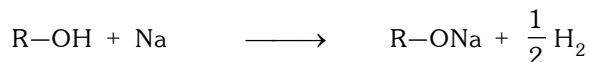
(i) Acidic nature :



Alcohols are less acidic than H_2O and neutral for litmus paper and gives H_2 with active metals (Na, K)

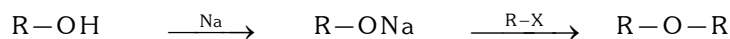
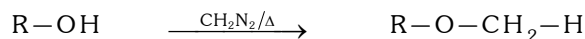


(ii) Reaction with CS_2 :



Sodium alkyl xanthate (Used as floating agent)

(iii) Alkylation:

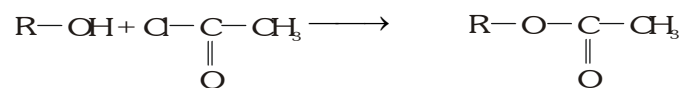


(Williamson synthesis)

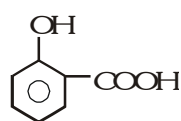
(iv) Acylation :



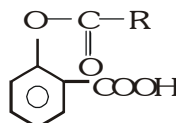
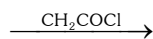
(Acylation)



(Acetylation)



Salicylic acid

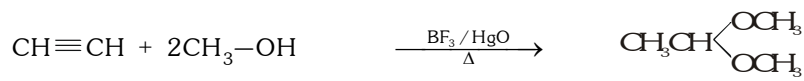


Acetoxy benzoic acid

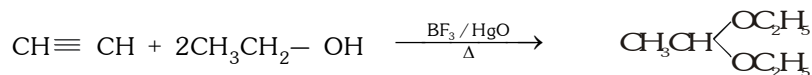
Acetyl salicylic acid

Aspirin [Used as analgesic]

(vii) Reaction with $\text{CH}\equiv\text{CH}$:



Methylal

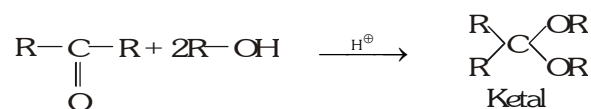


Ethylal

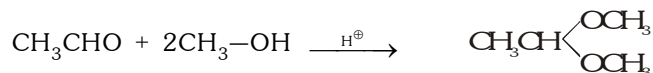
(viii) Reaction with carbonyl compounds :



Acetal

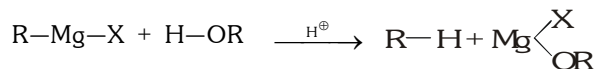


Ketal

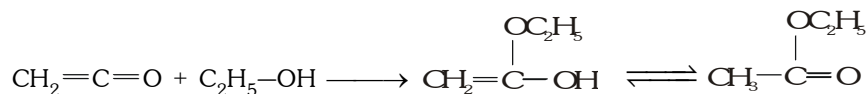
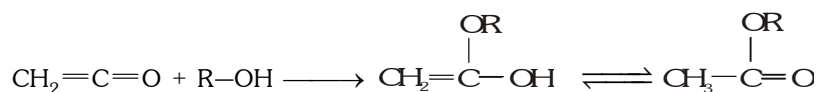


Methylal

(ix) Reaction with Grignard reagent :

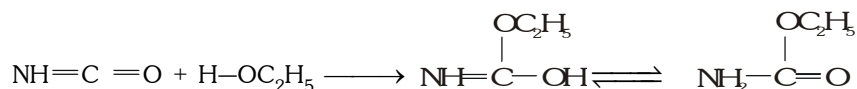


(x) Reaction with Ketene : Ketene is used as acetylating agent.



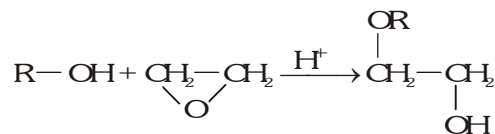
Ethylacetate

(xi) Reaction with isocyanic acid : Ethyl urethane is used in preparation of urea

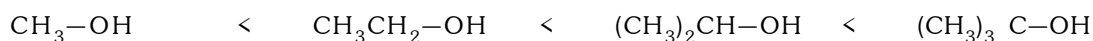


Ethyl urethane

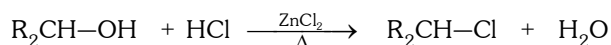
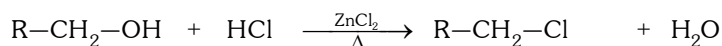
(xii) Reaction with oxirane :



(B) Reaction involving cleavage of $\text{C}-\text{OH}$: Reactivity order or basic nature is



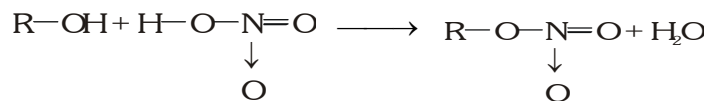
(i) **Reaction with halogen acid :**



2^o alcohol

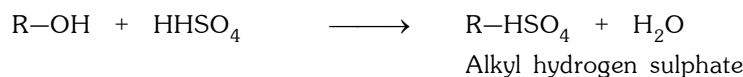
Reactivity of the acids is HI > HBr > HCl > HF

(ii) **Reaction with inorganic acids :**

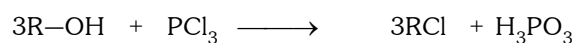


Nitric acid

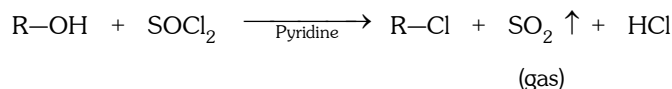
Alkyl nitrate



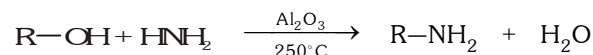
(iii) **Reaction with phosphorous halides :**



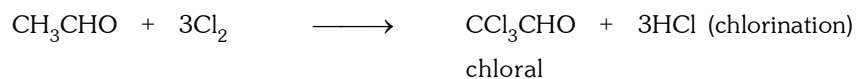
(iv) **Reaction with thionyl chloride (SOCl₂) :**



(v) **Reaction with NH₃ :** Alumina (Al₂O₃) is used as dehydrating agent.



(vi) **Reaction with halogens :** Oxidation and chlorination takes place simultaneously.

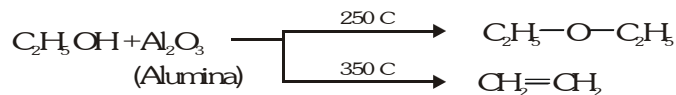
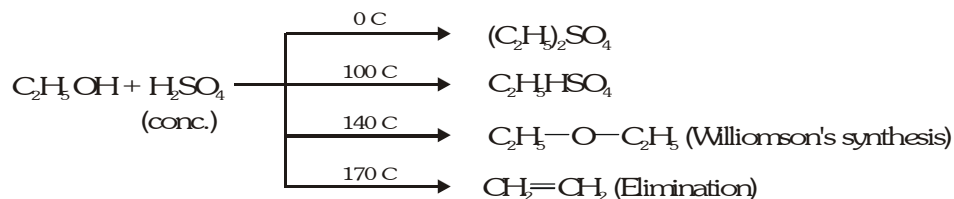


(C) **Reaction involving complete molecule of alcohol :**

(i) **Dehydration :** Removal of H₂O by two type

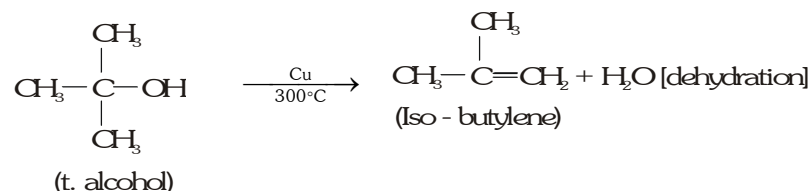
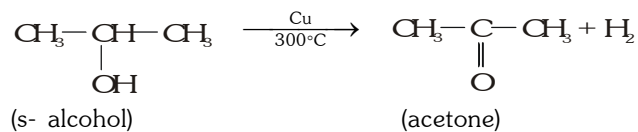
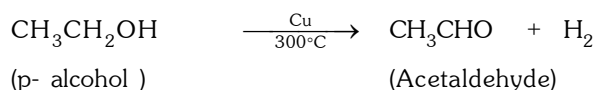
(a) Intermolecularly removal of H₂O [form ether]

(b) Intramolecularly removal of H₂O [form alkene]

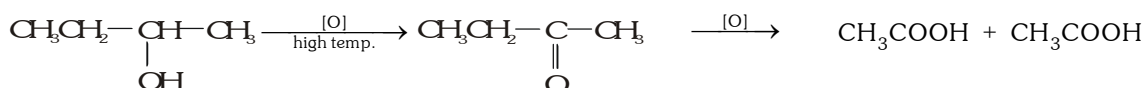
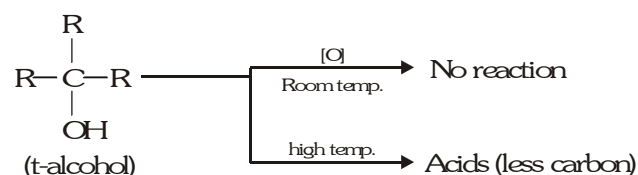
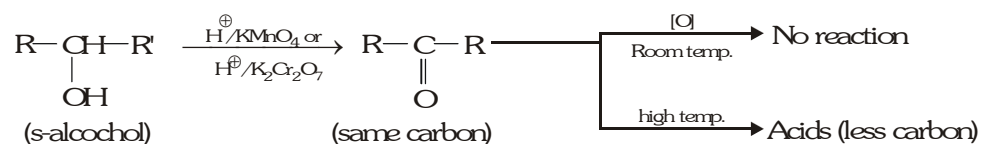
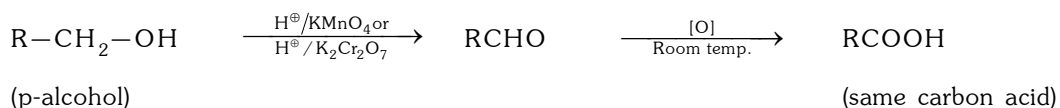


Ease of dehydration follow the order : **3 ROH > 2 ROH > 1 ROH > CH₃OH**

(ii) **Catalytic Dehydrogenation** : This reaction is useful in distinction of 1 , 2 and 3 alcohols.

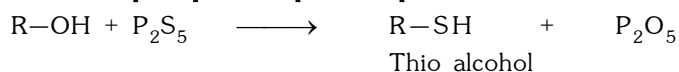


(iii) **Oxidation** : This reaction is useful in distinction of 1 , 2 and 3 alcohols.

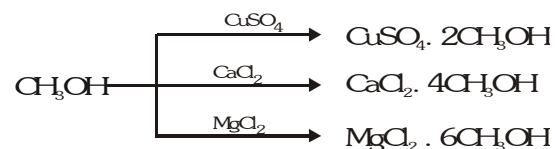


Carbonyl group goes with smaller alkyl group

(iv) **Reaction with phosphorous pentasulphide** :



(v) **Reaction with salts** :



(vi) **Distinction between 1 , 2 and 3 alcohols** :

(a) **Lucas test** : A mixture of HCl(conc.) and anhydrous ZnCl_2 is called Lucas reagent.

p-alcohol $\xrightarrow{\text{ZnCl}_2 + \text{HCl}}$ No turbidity at room temp. [On heating within 30 minutes.]

s-alcohol $\xrightarrow{\text{ZnCl}_2 + \text{HCl}}$ Turbidity appears within 5 minutes.

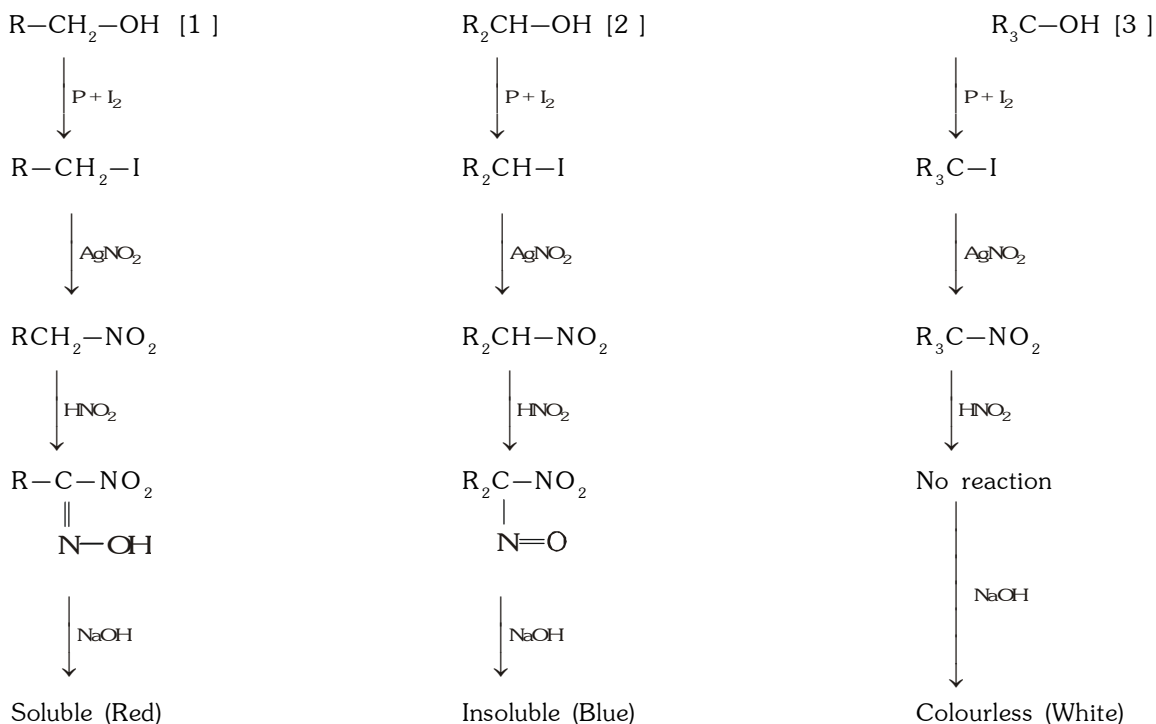
t-alcohol $\xrightarrow{\text{ZnCl}_2 + \text{HCl}}$ Turbidity appears within 1 minute.

(b) **Victor - Meyer test** : This is colour test for alcohol (pri. sec. & tert.) .

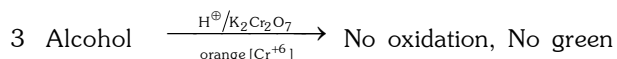
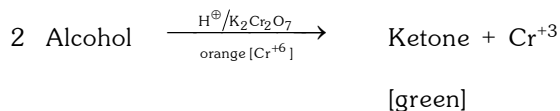
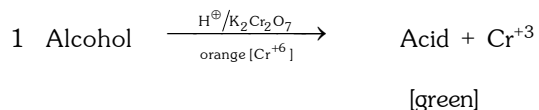
p-alcohol \longrightarrow Red colour

s-alcohol \longrightarrow Blue colour

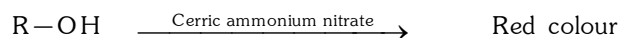
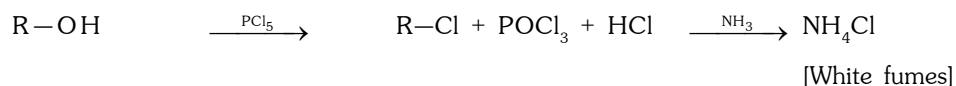
t-alcohol \longrightarrow No colour



(vii) **Dichromate test** :

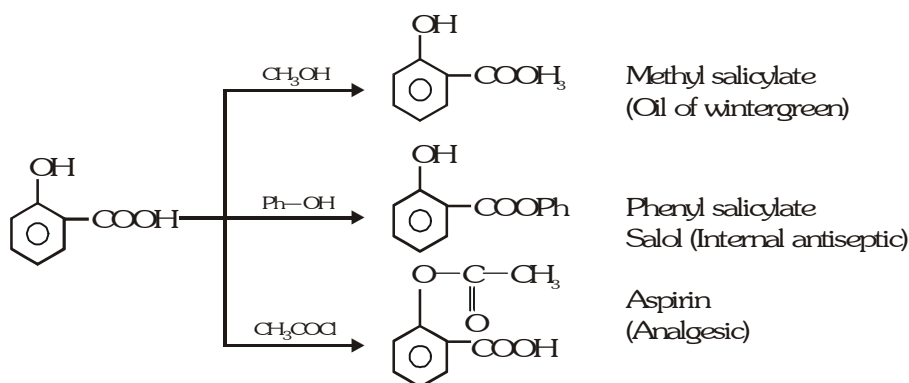


(viii) **Test of alcholic group** :



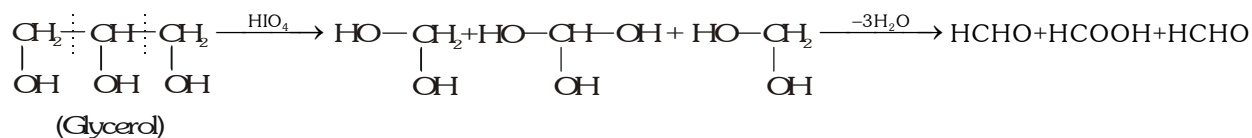
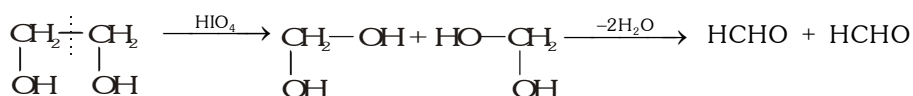
(ix) **Distinction between CH₃ - OH and C₂H₅OH**

	CH ₃ OH	CH ₃ CH ₂ OH
B.P.	65 C	78 C
I ₂ + NaOH	No ppt	Yellow ppt of CHI ₃
Cu/300 C	Smell of formalin [HCHO]	No smell
Salicylic acid	Smell of oil of wintergreen	No smell



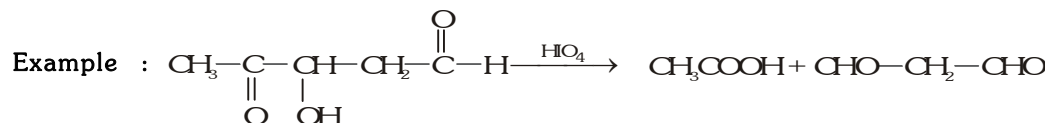
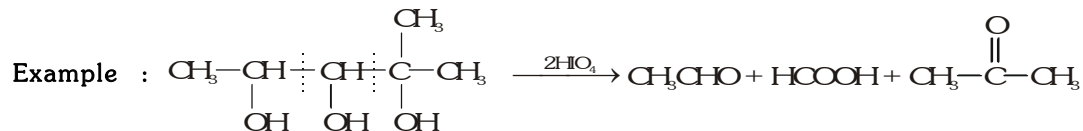
□ Additional reactions :

(a) Oxidation by HIO_4 [per iodic acid] :

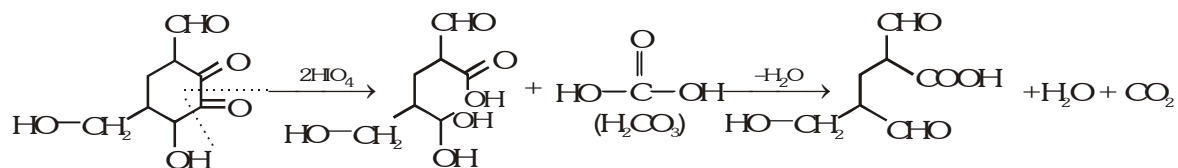


◆ Condition for oxidation by HIO_4 :

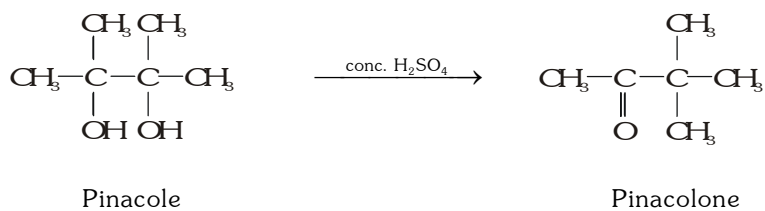
At least 2 -OH or 2 >C=O or 1 -OH and 1 >C=O should be at adjacent carbons.



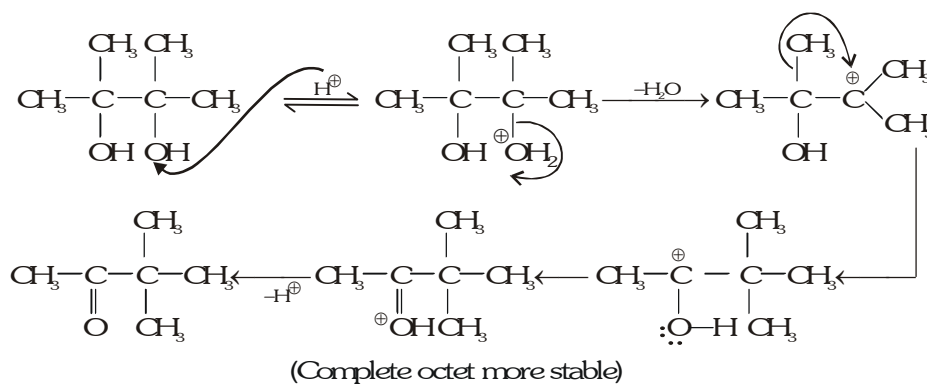
Example :



(b) Pinacole - Pinacolone Rearrangement :



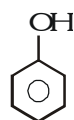
Mechanism :



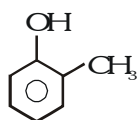
AROMATIC HYDROXY DERIVATIVES

□ Phenolic compounds :

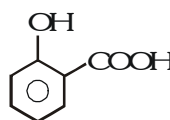
Compounds in which $-OH$ group is directly attached to sp^2c [Benzene ring]



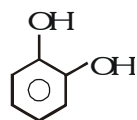
Phenol



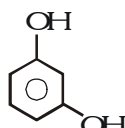
o-cresol



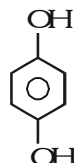
Salicylic acid



catechol

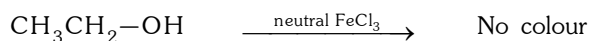
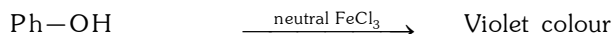


resorcinol



quinol

All phenolic compounds give characteristic colour with neutral $FeCl_3$.



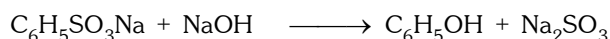
PHENOL (C_6H_5OH)

Phenol is also known as carbolic acid or Benzenol or hydroxy benzene. In phenol $-OH$ group is attached with sp^2 hybridised carbon. It was discovered by Runge in the middle oil fraction of coaltar distillation and named it carbolic acid (carbo = coal; oleum = oil). It is also present in traces in human urine.

□ General Methods of preparation :

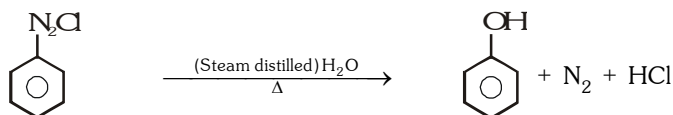
(1) From benzene sulphonic acid :

When sodium salt of benzene sulphonic acid is fused with $NaOH$ phenol is obtained.

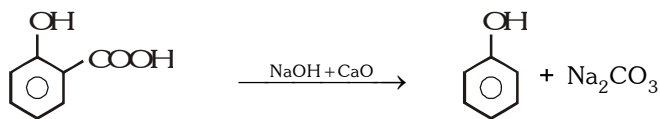


(2) From benzene diazonium chloride :

When benzene diazonium chloride solution is warmed, phenol is obtained with evolution of nitrogen.

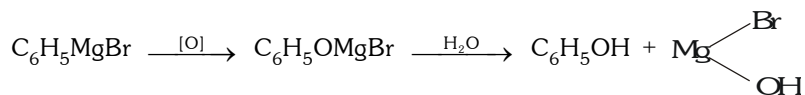


(3) By distilling a phenolic acid with sodalime (decarboxylation):

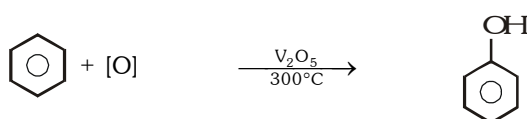


Salicylic acid

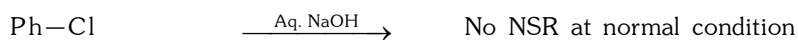
(4) From Grignard reagent : (The Grignard reagent on reaction with oxygen and subsequent hydrolysis by acid yields phenol)



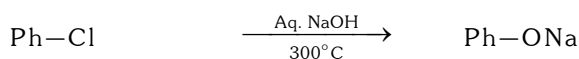
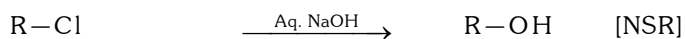
(5) From benzene :



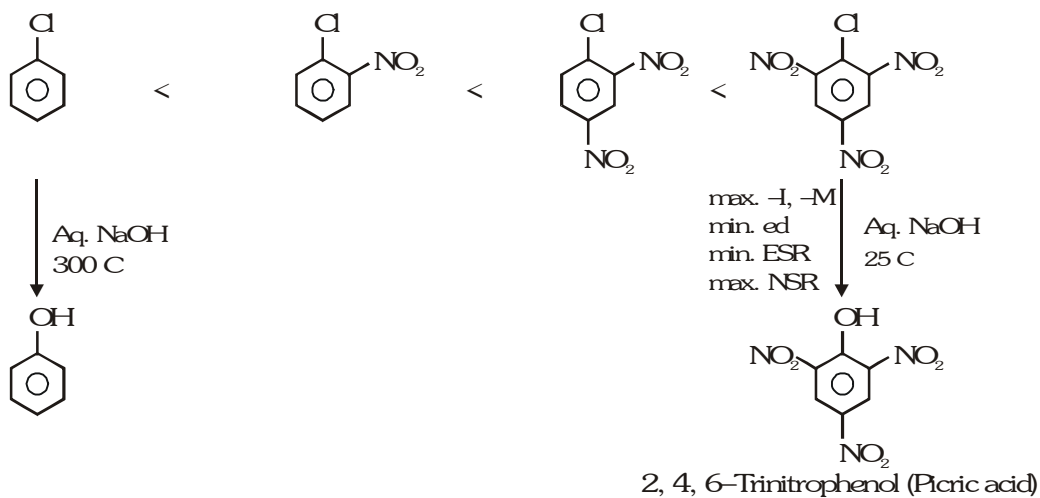
(6) From chloro benzene :



Stable by resonance



Order of NSR :

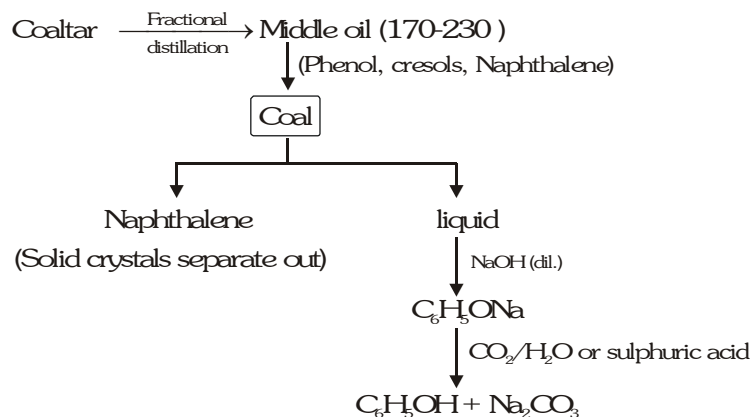


(7) Industrial preparation of phenol:

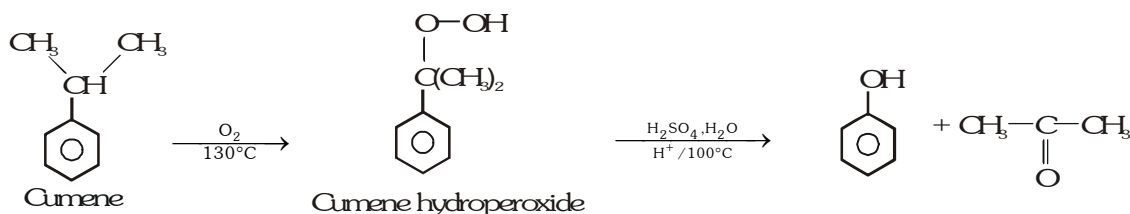
Phenol can be prepared commercially by :

- Middle oil fraction of coaltar distillation
- Cumene
- Raschig process
- Dow's process

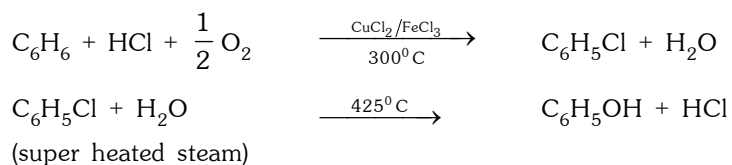
(a) Middle oil fraction of coaltar:



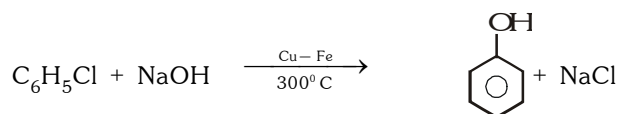
(b) **From cumene (Isopropyl benzene)** : Cumene is oxidised with oxygen into cumene hydroperoxide in presence of a catalyst. This is decomposed by dil. H_2SO_4 into phenol and acetone.



(c) **Raschig process** : Chlorobenzene is formed by the interaction of benzene, HCl and air at 300°C in presence of catalyst $\text{CuCl}_2 + \text{FeCl}_3$. It is hydrolysed by superheated steam at 425°C to form phenol and HCl.

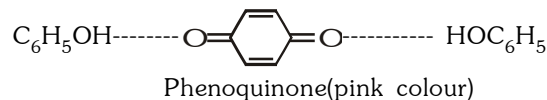


(d) **Dow process** : This process involves alkaline hydrolysis of chloro benzene-(large quantities of phenol formed).



□ **Physical properties :**

- (i) Phenol is a colourless, hygroscopic crystalline solid.
- (ii) It attains pink colour on exposure to air and light. (slow oxidation)

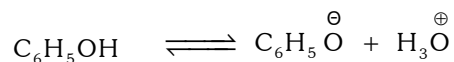


- (iii) It is poisonous in nature but acts as antiseptic and disinfectant.
- (iv) Phenol is slightly soluble in water, readily soluble in organic solvents.
- (v) Solubility of phenol in water is much lower than alcohols because of larger hydrocarbon part in the molecule.
- (vi) Due to intermolecular H-Bonding, phenol has relatively high boiling point than the corresponding hydrocarbons, aryl halides etc. but intermolecular H-bonding in o-derivatives is used in the preparation of dyes, drugs, bakelite and its melting point (MP) is 43°C and boiling point (BP) is 182°C .

□ **Chemical Properties :**

(A) **Reactions due to -OH group :**

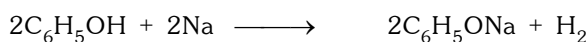
- ◆ **Acidic Nature :** Phenol is a weak acid. The acidic nature of phenol due to formation of stable phenoxide ion in solution. The phenoxide ion is stable due to resonance. The negative charge is spread through out the benzene ring which is stabilising factor in the phenoxide ion. Electron withdrawing groups (-NO₂, -Cl) increase the acidity of phenol while electron releasing groups (-CH₃ etc.) decrease the acidity of phenol.



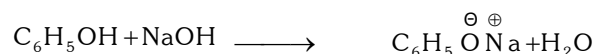
Phenol is stronger acid than alcohols but weaker than the carboxylic acids and even carbonic acid

The acidic nature of phenol is observed in the following:

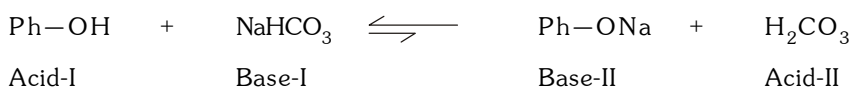
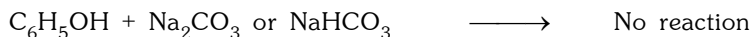
- (i) Phenol changes blue litmas to red.
 (ii) Highly electro positive metals react with phenol.



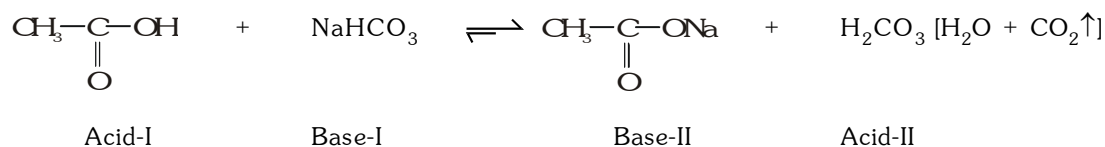
- (iii) Phenol reacts with strong alkalies to form phenoxides.



- (iv) However phenol does not decompose Na₂CO₃ or NaHCO₃ because phenol is weaker than carbonic acid.

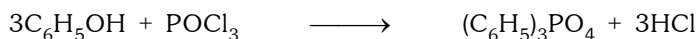
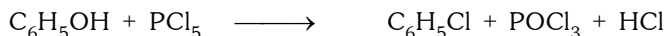


- (v) Phenol does not react with NaHCO₃.

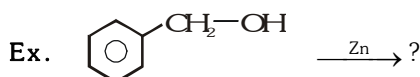
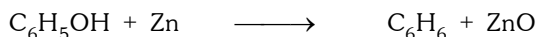


- (vi) Acetic acid reacts with NaHCO₃ and gives effervesence of CO₂.

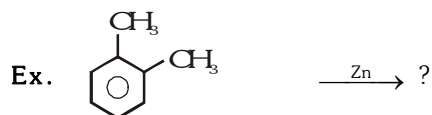
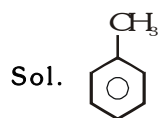
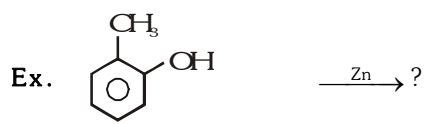
- ◆ **Reaction with PCl₅ :** Phenol reacts with PCl₅ to form chloro benzene. The yield of chlorobenzene is poor and mainly triphenyl phosphate is formed.



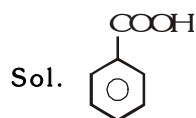
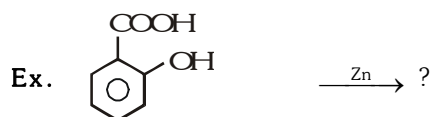
- ◆ **Reaction with Zn dust:** When phenol is distilled with zinc dust benzene is obtained.



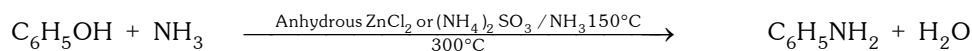
Sol. No reaction



Sol. No reaction



- ◆ **Reaction with NH₃ (Bucherer reaction):** Phenol reacts with NH₃ in presence of anhydrous ZnCl₂ to form aniline.

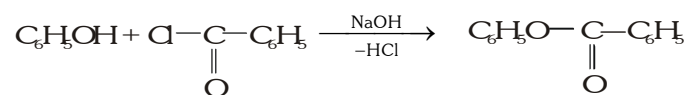
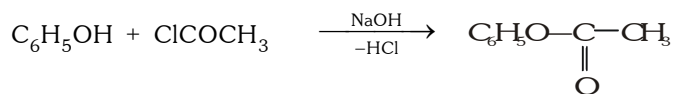


- ◆ **Reaction with FeCl₃:** Phenol gives violet colouration with FeCl₃ solution (neutral) due to formation of a complex.

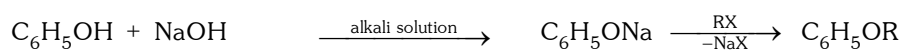


This reaction is used to differentiate phenol from alcohols.

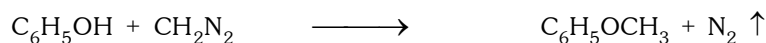
- ◆ **Acetylation (Schotten-Baumann reaction) :** Phenol reacts with acid chlorides or acid anhydrides in alkali solution to form phenyl esters.



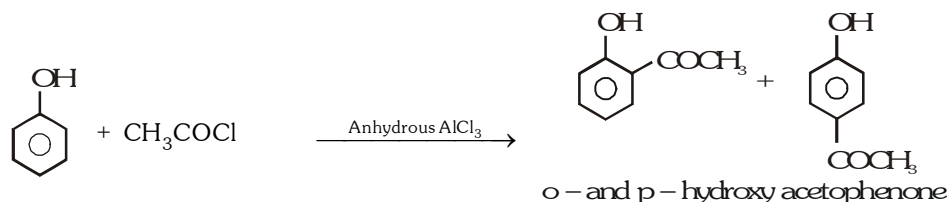
- ◆ **Ether formation (Alkylation) :** Phenol reacts with alkyl halides in alkali solution to form phenyl ethers. (Williamson's synthesis)



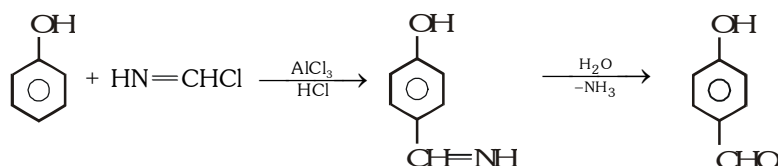
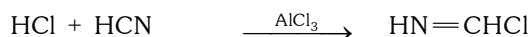
Sodium phenoxide



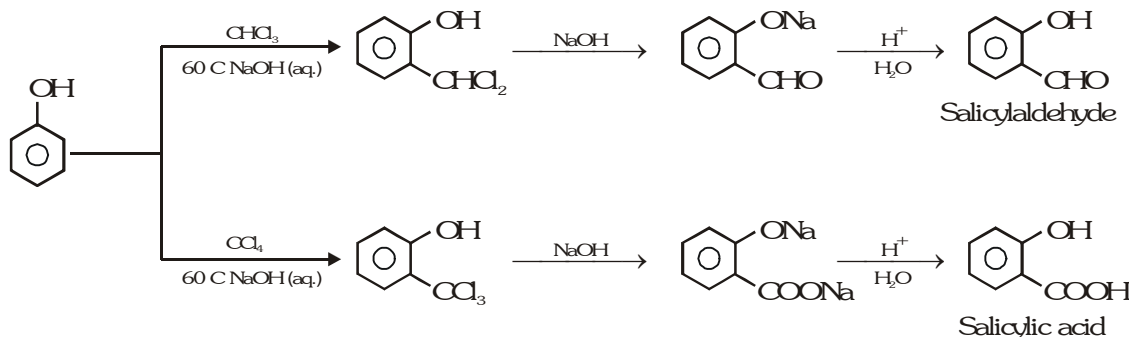
- ◆ **Reaction with P₂S₅ :** $5\text{C}_6\text{H}_5\text{OH} + \text{P}_2\text{S}_5 \xrightarrow{\Delta} 5\text{C}_6\text{H}_5\text{SH} + \text{P}_2\text{O}_5$



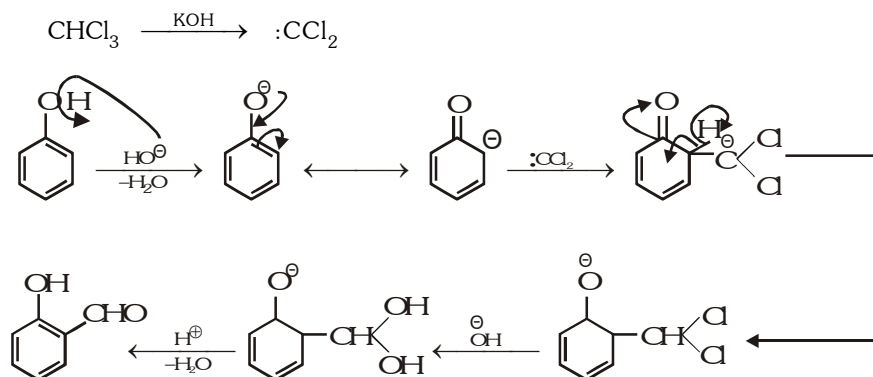
- ◆ **Gattermann aldehyde synthesis** : When phenol is treated with liquid HCN and HCl gas in presence of anhydrous AlCl_3 yields mainly p-hydroxy benzaldehyde (formylation)



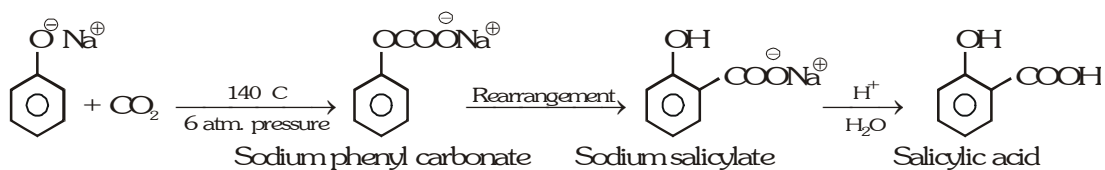
- ◆ **Riemer-Tiemann reaction** : Phenol on refluxing with chloroform and NaOH (aqueous) followed by acid hydrolysis yields o-hydroxy benzaldehyde. When CCl_4 is used salicylic acid is formed.



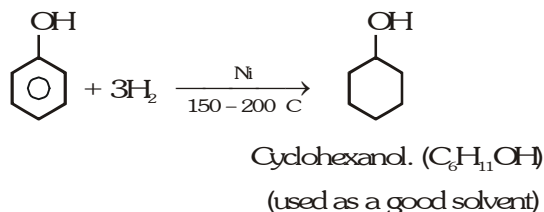
Mechanism : CCl_2 is neutral attacking electrophile (formed by α,α - elimination reaction)



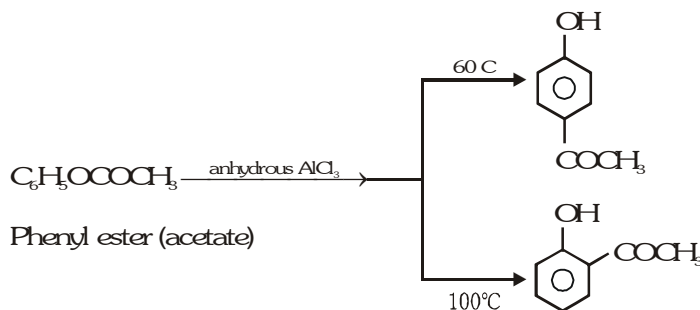
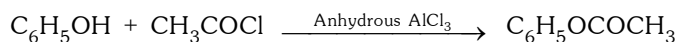
- ◆ **Kolbe's Schmidt reaction** : This involves the reaction of $\text{C}_6\text{H}_5\text{ONa}$ with CO_2 at 140°C followed by acid hydrolysis salicylic acid is formed followed.



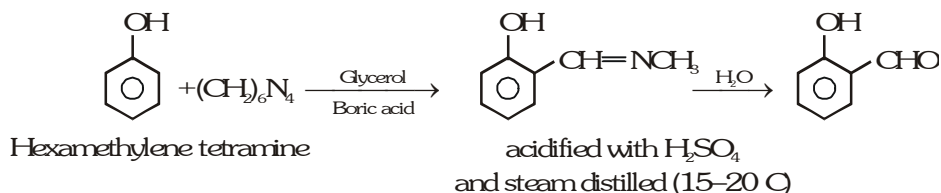
- ◆ **Hydrogenation:** Phenol when hydrogenated in presence of Ni at 150-200°C forms cyclohexanol.



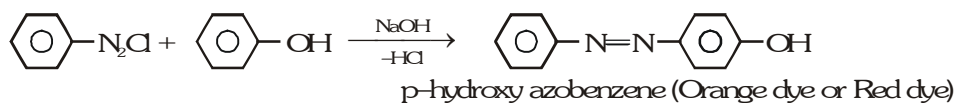
- ◆ **Fries rearrangement reaction :**



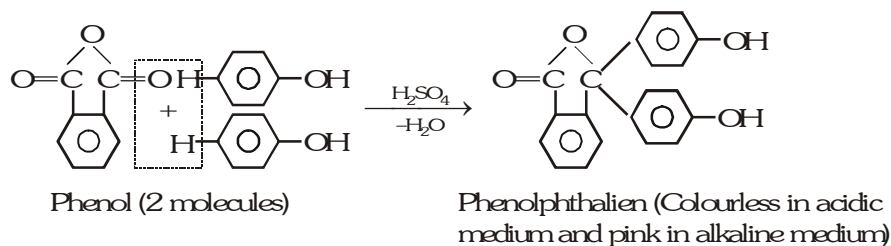
- ◆ **Duff's reaction:** This method gives only the o-compound which is hindered by the presence of a -I group in the ring.



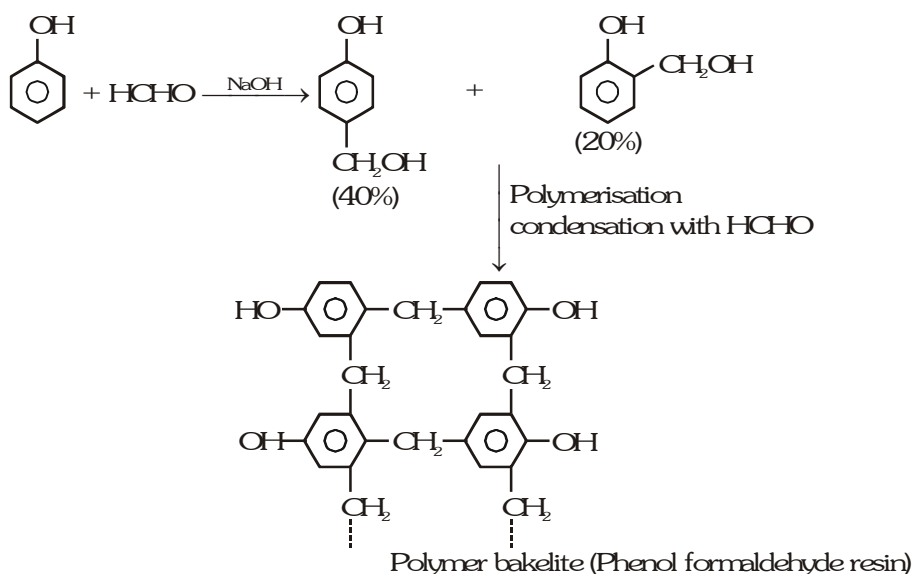
- ◆ **Coupling reactions:** Phenol couples with benzene diazonium chloride in presence of an alkaline solution to form a dye (p- hydroxy azobenzene) red only.



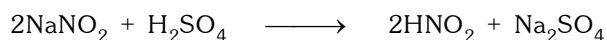
Phenol couples with phthalic anhydride in presence of conc. H₂SO₄ to form a dye (phenolphthalien) used as an indicator.



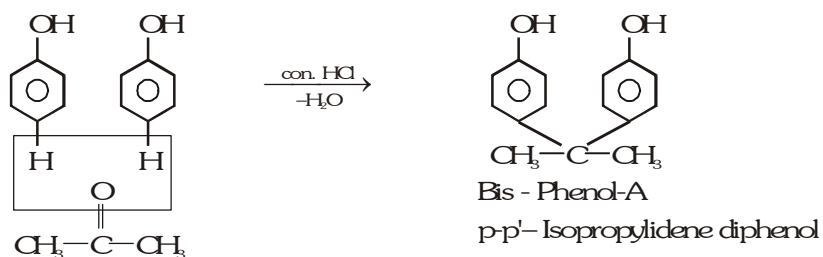
- ◆ **Lederer Manasse (Condensation with formaldehyde) :** Phenol condenses with HCHO (excess) in presence of NaOH or weak acid (H⁺) to form a polymer known as bakelite (aresin).



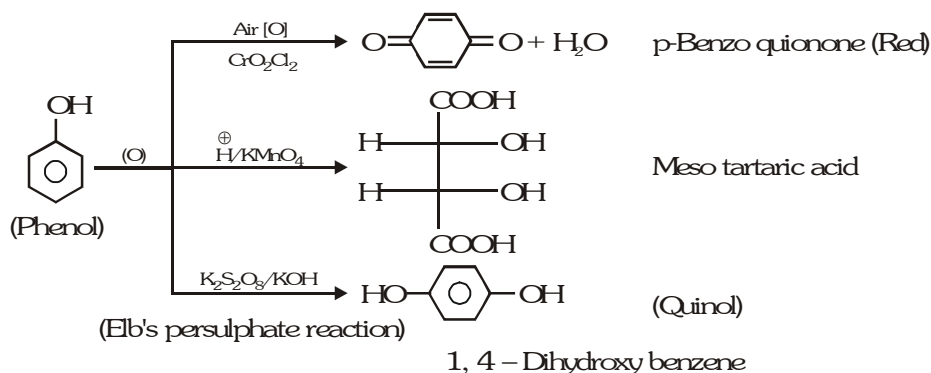
- ◆ **Leibermann's nitroso reaction** : When phenol is reacted with NaNO_2 and conc. H_2SO_4 it gives a deep green or blue colour which changes to red on dilution with water. When made alkaline with NaOH original green or blue colour is restored. This reaction is used as a test of phenol.



- ◆ **Reaction with acetone:** (Condensation with acetone)

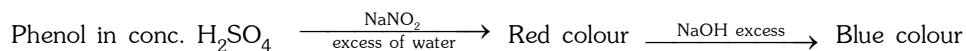


- ◆ **Oxidation:**



❑ **Test of Phenol :**

- (i) Phenol turns blue litmus to red.
- (ii) Aqueous solution of phenol gives a violet colour with a drop of ferric chloride.
- (iii) Phenol gives Lieber mann 's nitroso test.



- (iv) Aqueous solution of phenol gives a white ppt. of 2,4,6 tribromophenol with bromine water.
- (v) Phenol combines with phthalic anhydride in presence of conc. H_2SO_4 to form phenolphthalein which gives pink colour with alkali.
- (vi) With ammonia and sodium hypochlorite , phenol gives blue colour.

❑ **Differences between phenol and alcohol ($\text{C}_2\text{H}_5\text{OH}$) :**

- (i) Phenol is more acidic than aliphatic alcohol due to resonance in phenoxide ion.
- (ii) Phenol gives violet colour with FeCl_3 while aliphatic alcohol does not give.
- (iii) Phenol gives triphenyl phosphate with PCl_5 while aliphatic alcohol gives alkyl chloride.
- (iv) Phenol has phenolic odour whereas alcohol has pleasant odour.
- (v) Phenol on oxidation gives quinone while alcohol gives aldehyde or ketone and acids.

❑ **Uses of Phenol :** Phenol is used :

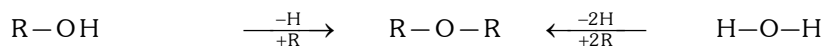
- (a) As an antiseptic in soaps and lotions. "Dettol" (2,4-Dichloro-3,5-dimethyl phenol)
- (b) In manufacture of azodyes, phenolphthalein , picric acid (explosive), cyclohexanol (Solvent for rubber), plastics (bakelite) etc.
- (c) In manufacture of drugs like aspirin salol, phenacetin etc.
- (d) As preservative for ink.

ETHER

R-O-R (Dialkyl ether), alkoxy alkane. Its General formula is $C_nH_{2n+2}O$.

$CH_3-O-CH_2CH_3$ (Methoxy ethane) or ethyl methyl ether or 2-oxa butane

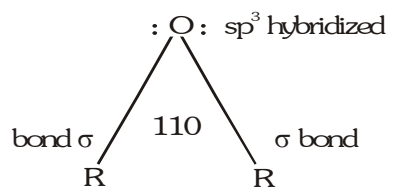
Ether is monoalkyl derivative of R-OH and dialkyl derivative of H_2O



Classification : They may be classified as :

- (a) Simple or symmetrical ether. e.g. R-O-R
- (b) Mixed or unsymmetrical ether e.g. R-O-R'

Structure :

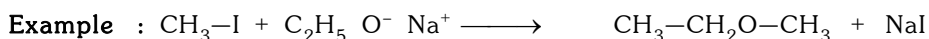


The molecule of ether is bent due to lone pair of electron on oxygen atom- bond electron repulsion. The bond angle is 110° . It is greater than that of water 105° due to the repulsion between bulky alkyl groups. Due to bent structure, it posses dipole moment and hence are polar molecules.

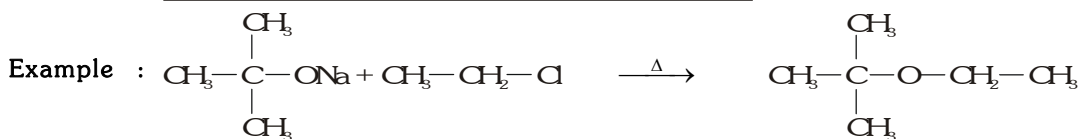
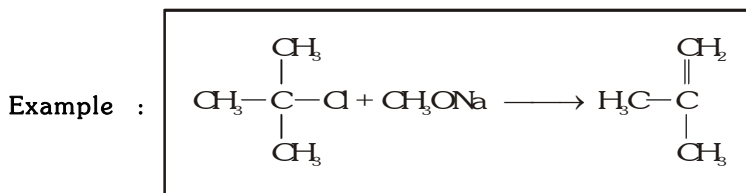
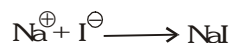
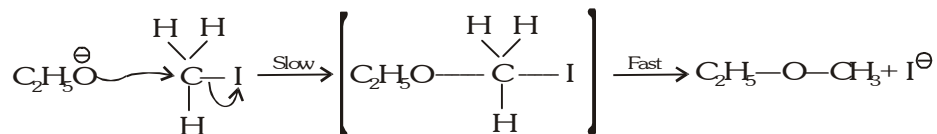
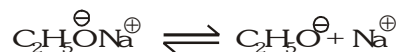
□ **General Methods of Preparation :**

(A) **From alkyl halides :**

(i) **By Williamson's synthesis :**

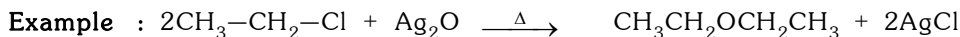


Mechanism : [S_N² Reaction]



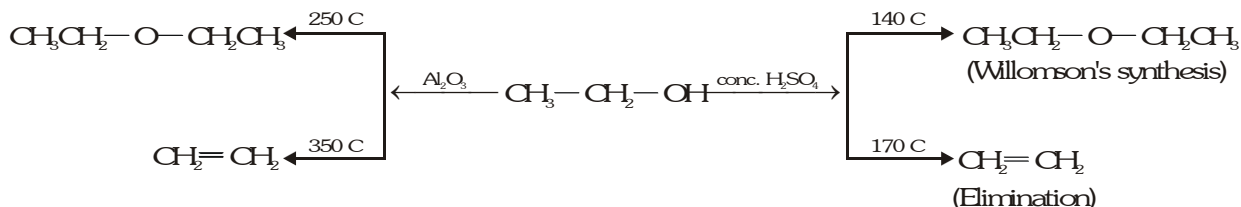
[Stable by Resonance]

(ii) Reaction with Dry Ag₂O :

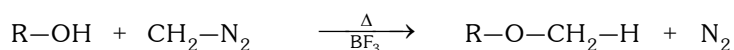


(B) From R-OH:

(i) By dehydration : $R-OH \xrightarrow[\Delta]{Con. H_2SO_4} ?$



(ii) Reaction with CH₂N₂ (diazomethane) :



Physical Properties :

- (i) CH₃OCH₃ , CH₃OCH₂CH₃ are gases and higher are volatile liquids.
- (ii) Ether are less polar [$\mu=1.18D$].
- (iii) Ethers are less soluble in H₂O.
- (iv) Ethers have less BP then corresponding alcohol.

Ex. Ethers are less soluble in H₂O . Why ?

Sol. Reason : Due to less polar, it forms weaker H-Bonding with H₂O.

Ex. Ethers have less BP then corresponding alcohol. Why ?

Sol. Reason : No H-Bonding in ether molecules.

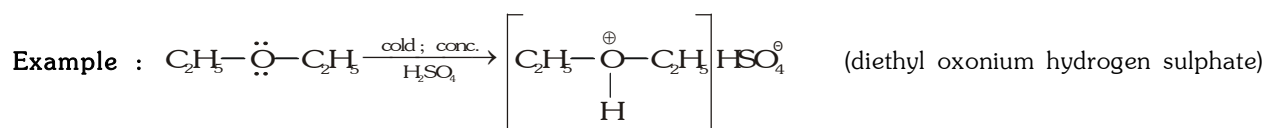
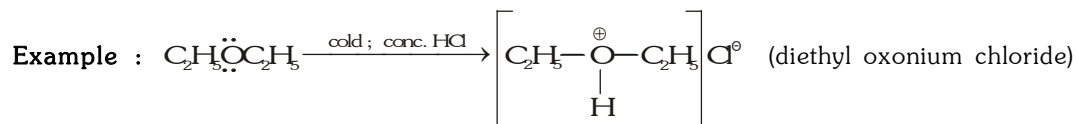
Chemical properties :

Ethers are less polar so less reactive and do not react with active metals [Na,K], cold dil. acid, oxidising and reducing agent.

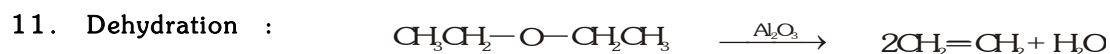
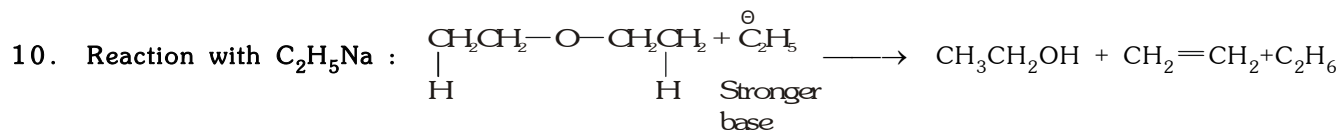
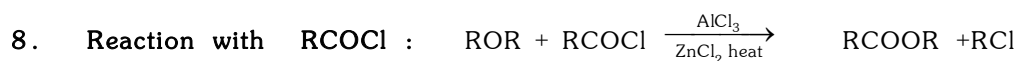
Reason : They do not have any active functional group.

1. **Basic nature** : Due to presence of *l.p* on oxygen atom ether behave as lewis base

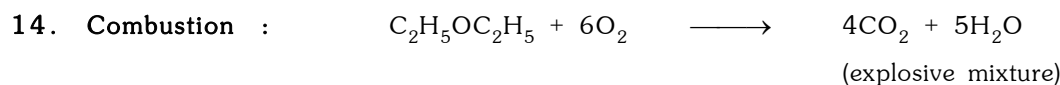
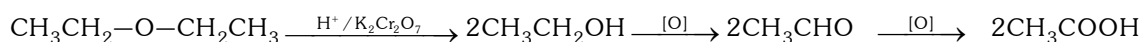
Ethers react with cold conc. acid and form oxonium ion



Ether form dative bond with Lewis acids like BF₃, AlCl₃, RMgX etc.



13. Oxidation :

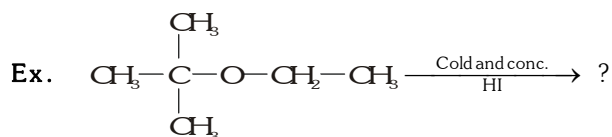


15. Reaction with HX : Reactivity of HX $\text{HI} > \text{HBr} > \text{HCl}$

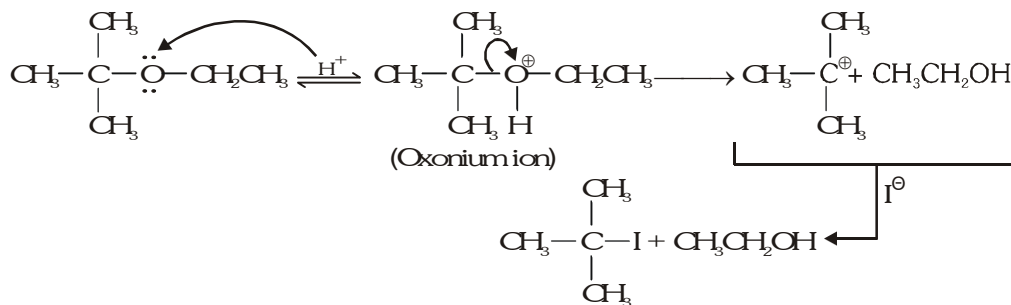
□ Reaction with cold conc. HX :

Ethers forms oxonium salt with cold and conc. HCl (less reactive)

Cold conc. HI and HBr (more reactive) break C-O bond.

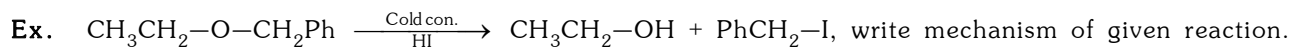


Sol. Mechanism

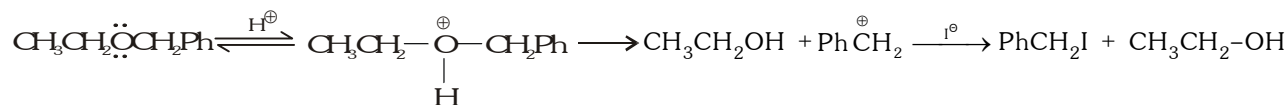


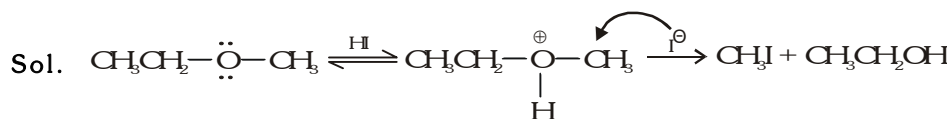
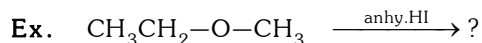
◆ If oxonium ion gives more stable carbocation [PhCH_2^+ , $\text{CH}_2=\text{CH}-\overset{\oplus}{\text{C}}\text{H}_2$, $(\text{CH}_3)_3\overset{\oplus}{\text{C}}$] then SN^1 reaction occurs.

◆ If oxonium ion gives less stable carbocation [Ph^+ , $\text{CH}_2=\overset{\oplus}{\text{C}}\text{H}$, $\text{CH}_3\overset{\oplus}{\text{C}}\text{H}_2$] then SN^2 reaction occurs, and X^\ominus attacks at less hindered carbon.



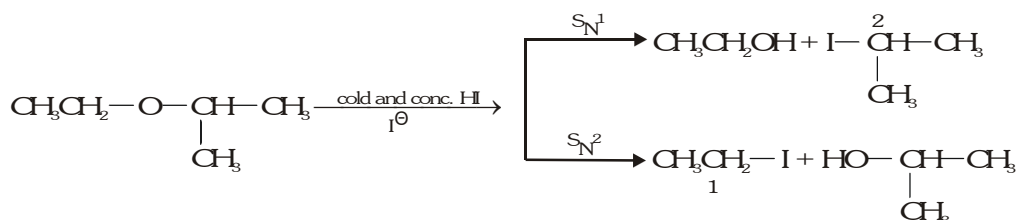
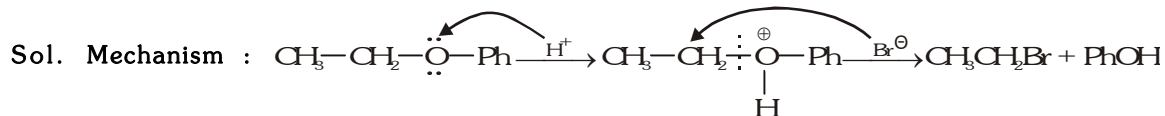
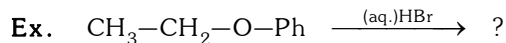
Sol. Mechanism :



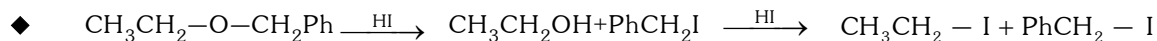


Oxonium ion gives less stable carbocation

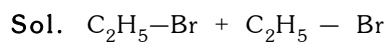
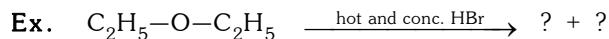
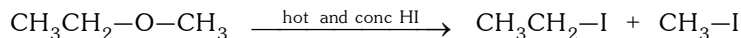
S_N^2 reaction I^\ominus attacks at less hindered carbon.



◆ If excess of HI is used then two moles of alkyl halides are formed.



(B) Reaction with hot and conc. HX :



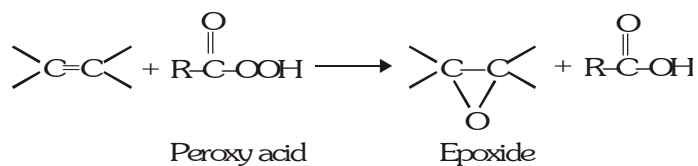
□ Uses of ether :

- (i) General anaesthetics agent.
- (ii) Refrigerant a mixture of ether and dry ice gives temperature as low 110 C.
- (iii) Solvent for oil, fats, resins, Grignard reagent.
- (iv) For providing inert & moist free medium to organic reaction example : Wurtz reactions.
- (v) In perfumery.
- (vi) Di-isopropyl ether \longrightarrow Petrol as an antiknock comp.
- (vii) Mixture of alcohol and ether is used as a substitute of petrol. Trade name "Natalite"
- (viii) Halothane (CF_3CHClBr) used as an anaesthetic because it produces unconsciousness without affecting lung and heat.

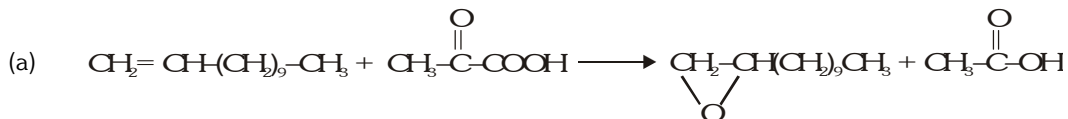
□ Preparation of Epoxides :

- (i) Epoxidation of alkenes by reaction with peroxy acids
- (ii) Base-promoted ring closure of vicinal halohydrins
- (iii) Epoxidation of alkenes by reaction with peroxy acids

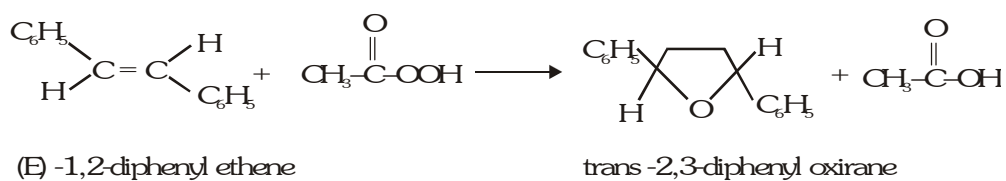
◆ **Epoxydation of alkenes by reaction with peroxy acids :**



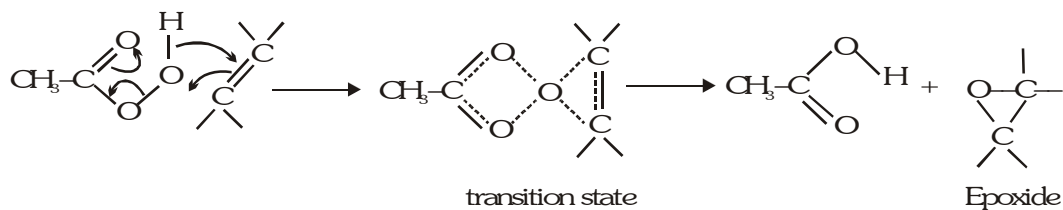
Example :



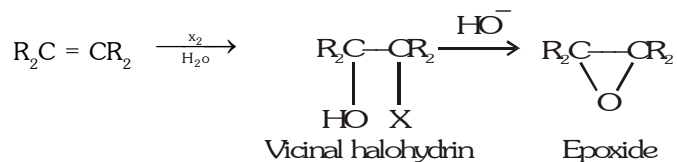
(c) Epoxydation is a stereospecific syn addition :



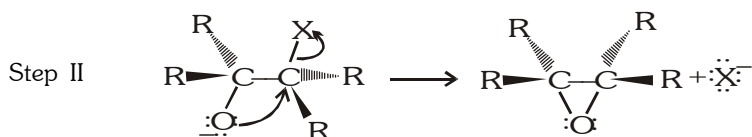
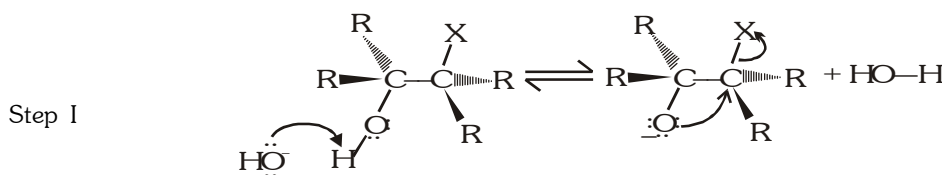
Mechanism :



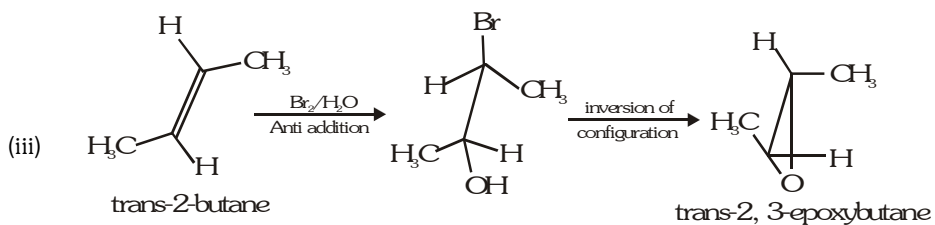
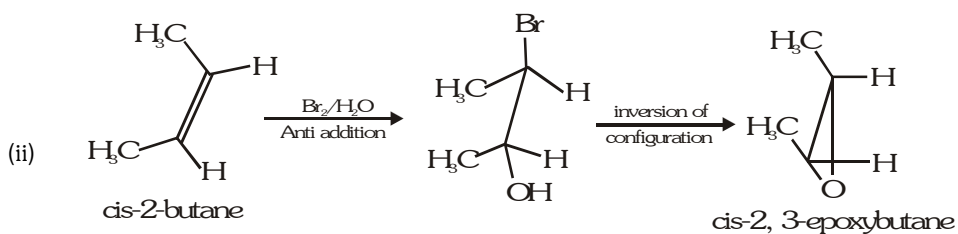
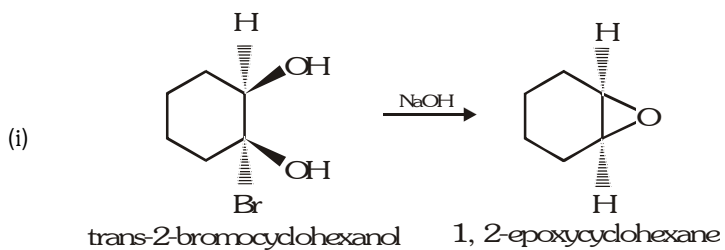
◆ **Base-promoted ring closure of vicinal halohydrins :**



Mechanism :

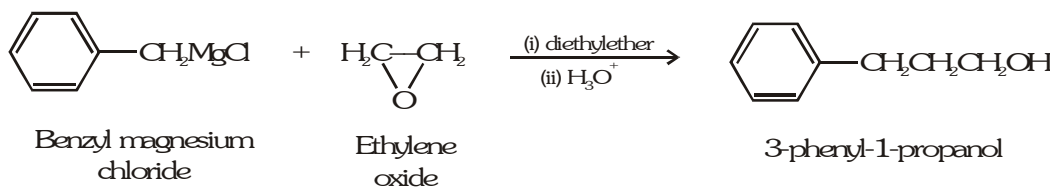
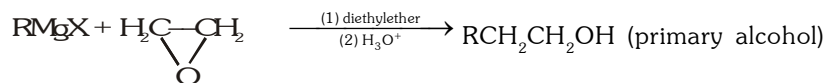


Example :

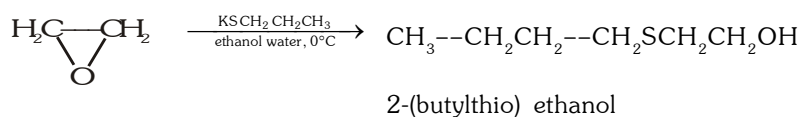
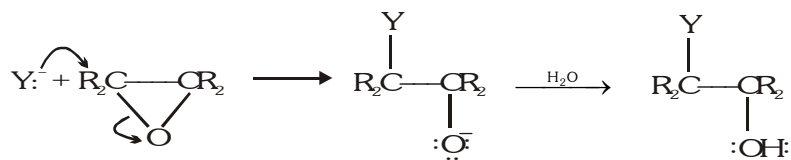


◆ Reaction of Epoxides :

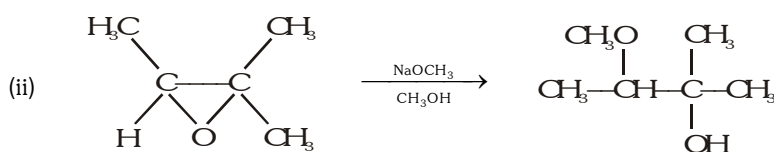
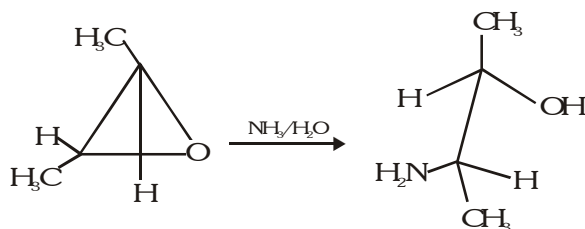
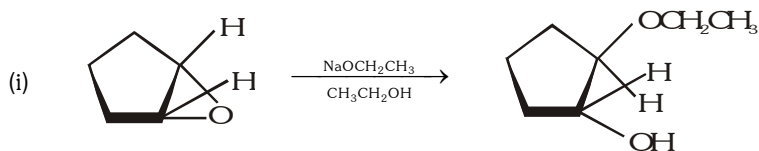
◆ With Grignard reagent :



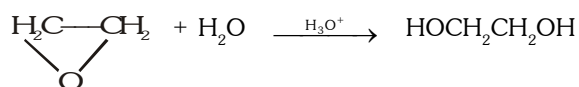
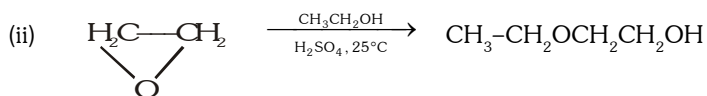
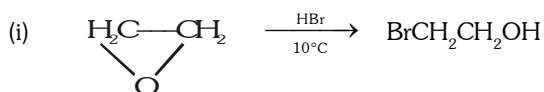
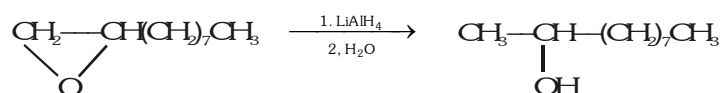
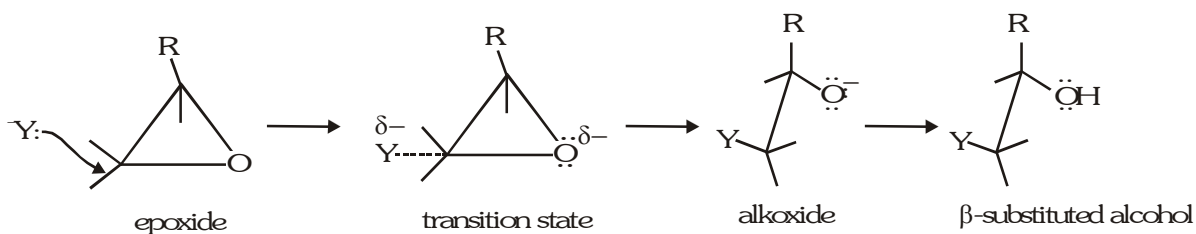
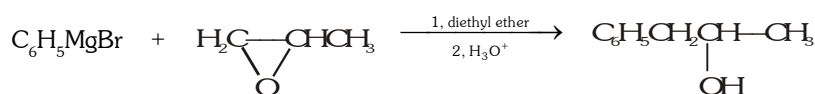
◆ Nucleophilic ring opening reactions of epoxides :



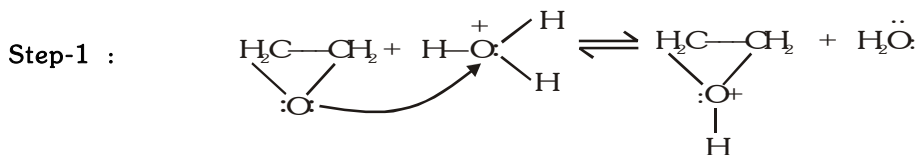
Note : Nucleophilic ring opening reactions of epoxides is the characteristic feature of S_N2 reaction.

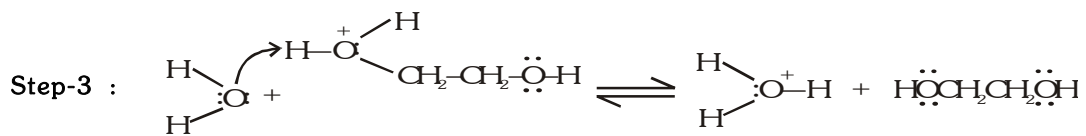
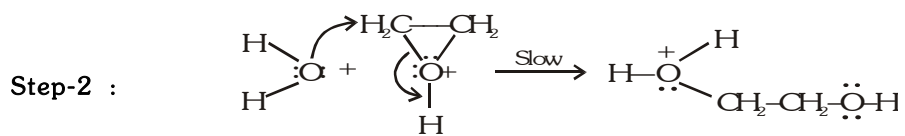


◆ **Nucleophilic ring opening of epoxides :**



Mechanism





Example :

